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to Charles's law, when the temperature of a sample of gas in-  
creases, the volume of the gas increases by the same factor.  
Therefore, doubling the Kelvin temperature of a gas will double its  
volume. Reduc-ing the Kelvin temperature by 25% will reduce the  
volume by 25%.

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pressure nor temperature is constant; therefore, the combined gas law must be used to calculate the volume of CO<sub>2</sub> at the final temperature and pressure. multiply by the volume ratio, CO<sub>2</sub> use the combined gas law to convert from the initial temperature and pressure to the final temperature and pressure C<sub>2</sub>H<sub>4</sub>

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as accurately as you can, gas you record the following data: Rules for Determining Significant Figures A. Mass of solution in g rearrange  $m D V$  to solve for  $V m$ .

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